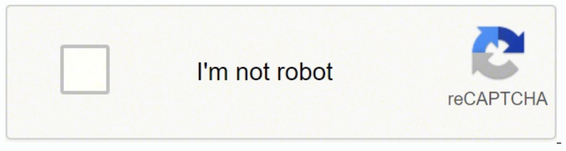

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Balancing act worksheet chemistry

Balancing Equations Practice Worksheet

Balance the following equations:

1) ___ $NaNO_3$ + ___ PbO → ___ $Pb(NO_3)_2$ + ___ Na_2O
2) ___ AgI + ___ $Fe_2(CO_3)_3$ → ___ FeI_3 + ___ Ag_2CO_3
3) ___ $C_2H_2O_2$ + ___ O_2 → ___ CO_2 + ___ H_2O
4) ___ $ZnSO_4$ + ___ Li_2CO_3 → ___ $ZnCO_3$ + ___ Li_2SO_4
5) ___ V_2O_5 + ___ CaS → ___ CaO + ___ V_2S_5
6) ___ $Mn(NO_3)_4$ + ___ $BeCl_2$ → ___ $Be(NO_3)_2$ + ___ $MnCl_4$
7) ___ AgBr + ___ $GaPO_4$ → ___ Ag_3PO_4 + ___ $GaBr_3$
8) ___ H_2SO_3 + ___ $Ni(OH)_3$ → ___ $Ni_2(SO_3)_3$ + ___ H_2O
9) ___ S_8 + ___ O_2 → ___ SO_2
10) ___ Fe + ___ $AgNO_3$ → ___ $Fe(NO_3)_2$ + ___ Ag

Solutions for the Balancing Equations Practice Worksheet

1) $2\ NaNO_3 + PbO \rightarrow Pb(NO_3)_2 + Na_2O$
2) $6\ AgI + Fe_2(CO_3)_3 \rightarrow 2\ FeI_3 + 3\ Ag_2CO_3$
3) $C_2H_2O_2 + 2\ O_2 \rightarrow 2\ CO_2 + 2\ H_2O$
4) $ZnSO_4 + Li_2CO_3 \rightarrow ZnCO_3 + Li_2SO_4$
5) $V_2O_5 + 5\ CaS \rightarrow 5\ CaO + V_2S_5$

Balancing Equations

Name _____

Chem Worksheet 10-2

Balance the following equations by adding coefficients before the elements or compounds in the following equations.

1. $\_\_NH_4NO_3 \rightarrow \_\_N_2O + \_\_H_2O$
2. $\_\_K + \_\_H_2O \rightarrow \_\_H_2 + \_\_KOH$
3. $\_\_O_2 \rightarrow \_\_O_3$
4. $\_\_Mg + \_\_HCl \rightarrow \_\_MgCl_2 + \_\_H_2$
5. $\_\_SO_3 \rightarrow \_\_SO_2 + \_\_O_2$
6. $\_\_Cu + \_\_S_8 \rightarrow \_\_Cu_2S$
7. $\_\_Na_2O + \_\_H_2O \rightarrow \_\_NaOH$
8. $\_\_PCl_5 + \_\_H_2O \rightarrow \_\_H_3PO_4 + \_\_HCl$
9. $\_\_FeO + \_\_O_2 \rightarrow \_\_Fe_2O_3$
10. $\_\_NaOH + \_\_H_2SO_4 \rightarrow \_\_Na_2SO_4 + \_\_H_2O$
11. $\_\_Al + \_\_Cl_2 \rightarrow \_\_AlCl_3$
12. $\_\_NH_3 + \_\_Na \rightarrow \_\_H_2 + \_\_NaNH_2$
13. $\_\_C_9H_{20} + \_\_O_2 \rightarrow \_\_CO_2 + \_\_H_2O$
14. $\_\_CaSiO_3 + \_\_HF \rightarrow \_\_SiF_4 + \_\_CaF_2 + \_\_H_2O$

Worksheet – Balancing Equations

Ch 10.1 Std 3a

Name:_____
Period: _____
Date: _____

Directions: Fill in the blanks to balance each of the following equations. If coefficient is one, you must put the #1 in the blank. Blanks to left of the number of the equation will be completed as another assignment tomorrow.

____1. ___ $Pb(NO_3)_2$ + ___ HCl ---> ___ $PbCl_2$ + ___ HNO_3

____2. ___ Ca + ___ HCl ---> ___ $CaCl_2$ + ___ H_2

____3. ___ $MgCl_2$ + ___ NaOH ---> ___ $Mg(OH)_2$ + ___ NaCl

____4. ___ $Ca(OH)_2$ + ___ H_3PO_4 ---> ___ $Ca_3(PO_4)_2$ + ___ H_2O

____5. ___ K + ___ O_2 ---> ___ K_2O

____6. ___ $MgBr_2$ + ___ Cl_2 ---> ___ $MgCl_2$ + ___ Br_2

____7. ___ MgO + heat ---> ___ Mg + ___ O_2

____8. ___ Na + ___ HOH ---> ___ NaOH + ___ H_2

____9. ___ $Cu(NO_3)_2$ + ___ Zn ---> ___ $Zn(NO_3)_2$ + ___ Cu

____10. ___ K_2CO_3 + ___ $CaCl_2$ ---> ___ $CaCO_3$ + ___ KCl

____11. ___ H_2CO_3 ---> ___ H_2O + ___ CO_2

____12. ___ $Mg(ClO_3)_2$ ---> ___ $MgCl_2$ + ___ O_2

Balancing Chemical Equations

Balance the equations below:

1) ____ N_2 + ____ H_2 → ____ NH_3

2) ____ $KClO_3$ → ____ KCl + ____ O_2

3) ____ NaCl + ____ F_2 → ____ NaF + ____ Cl_2

4) ____ H_2 + ____ O_2 → ____ H_2O

5) ____ $Pb(OH)_2$ + ____ HCl → ____ H_2O + ____ $PbCl_2$

6) ____ $AlBr_3$ + ____ K_2SO_4 → ____ KBr + ____ $Al_2(SO_4)_3$

7) ____ CH_4 + ____ O_2 → ____ CO_2 + ____ H_2O

8) ____ C_3H_8 + ____ O_2 → ____ CO_2 + ____ H_2O

9) ____ C_8H_{18} + ____ O_2 → ____ CO_2 + ____ H_2O

10) ____ $FeCl_3$ + ____ NaOH → ____ $Fe(OH)_3$ + ____NaCl

11) ____ P + ____O_2 → ____P_2O_5

12) ____ Na + ____ H_2O → ____ NaOH + ____H_2

13) ____ Ag_2O → ____ Ag + ____O_2

14) ____ S_8 + ____O_2 → ____ SO_3

15) ____ CO_2 + ____ H_2O → ____ $C_6H_{12}O_6$ + ____O_2

16) ____ K + ____ MgBr → ____ KBr + ____ Mg

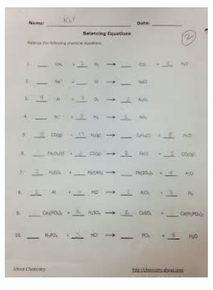
17) ____ HCl + ____ $CaCO_3$ → ____ $CaCl_2$ + ____H_2O + ____ CO_2

18) ____ HNO_3 + ____ $NaHCO_3$ → ____ $NaNO_3$ + ____ H_2O + ____ CO_2

19) ____ H_2O + ____ O_2 → ____ H_2O_2

20) ____ NaBr + ____ CaF_2 → ____ NaF + ____ $CaBr_2$

21) ____ H_2SO_4 + ____ $NaNO_2$ → ____ HNO_2 + ____ Na_2SO_4



This allows to extract from the extracellular liquid the material needed by the cell. 4: 4, 8, 12, 16 2: 2, 4, 6, 8 3. Multiply the numerator and denominator of each fragrance by the factor you would take to reach your common denominator. The most direct ways of membrane transport are passive. 4. If the cells of the plant become hypertonic, as occurs in the drought or if a plant is not adequately irrigated, the water will output from the cell. For Example, Think About Someone Opening A Bottle of Perfume in Room Filled With People. Concentration Gradient: An Area of High Concentration Across from An Area of Low Concentration Diffusion: A Passive Process of Transport Of Low-Molecular Weight DinW Its Concentration Gradient Facilitated Transport: A Process by Which Movie Movies Down A Concentration Gradient (from High To Low Concentration) Using Integral Membrane Proteins Hypertonic: Describe to Solution In Which Extracellular Fluid Higher Osmolarity Than The Fluid Inside The Cell Hypotonic: Describe to Solution In Which Extracellular Fluid Has Lower Osmolarity Than The Fluid Inside The Cell Isotonic: Describe to Solution in Which The Extracellular Fluid Has The Same Osmolarity as The Fluid Inside The Cell Osmolarity: The Total Amount Of Substances Dissolved In A Specific Amount Of Solution Osmosis: The Transport Of Water Through A Semipermeable Membrane from an Area of High Water Concentration To An Area of Low Water Concentration ACR Figure 3.27 The Turgency Pressure within a vegetable cell It depends on the tonicity of the solution in which it is batting.

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